

# CHEMISTRY

# COMPUTING FORMULA MASS WORKSHEET

## Directions:

Find the formula mass of the following compounds. Round atomic masses to the tenth of a decimal place. Place your final answer in the FORMULA MASS COLUMN.

Problem Set-up example:	
Find the formula mass of $\text{Ca}(\text{NO}_3)_2$	
Ca:	$1 \times 40.1 = 40.1$
N:	$2 \times 14.0 = 28.0$
O:	$6 \times 16.0 = 96.0$
Formula Mass =	$\underline{\quad 164.1 \quad}$

COMPOUND	FORMULA MASS
$\text{AgNO}_2$	
$\text{NiSO}_3$	
$\text{Ca}_3(\text{PO}_4)_2$	
$\text{HgSO}_4$	
$\text{Fe}(\text{NO}_3)_3$	
$\text{KBr}$	
$\text{BeCr}_2\text{O}_7$	
$\text{Co}(\text{ClO}_3)_2$	
$\text{Cu}_2\text{C}_4\text{H}_4\text{O}_6$	
$\text{CuSO}_4 \cdot 7 \text{H}_2\text{O}$	

COMPOUND	FORMULA MASS
$\text{ZnCl}_2$	
$\text{K}_3\text{PO}_4$	
$\text{Al}_2(\text{SO}_4)_3$	
$\text{MgCrO}_4$	
$\text{CaC}_4\text{H}_4\text{O}_6$	
$\text{NaCl}$	
$\text{K}_2\text{Cr}_2\text{O}_7$	
$\text{H}_2\text{SO}_4$	
$\text{Cu}(\text{OH})_2$	
$\text{MgSO}_4 \cdot 5 \text{H}_2\text{O}$	

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N:	$2 \times 14.0 = 28.0$
O:	$6 \times 16.0 = 96.0$
Formula Mass =	$\overline{164.1}$

COMPOUND	FORMULA MASS
$\text{AgNO}_2$	<b>153.9</b>
$\text{NiSO}_3$	<b>138.8</b>
$\text{Ca}_3(\text{PO}_4)_2$	<b>310.2</b>
$\text{HgSO}_4$	<b>296.7</b>
$\text{Fe}(\text{NO}_3)_3$	<b>241.9</b>
$\text{KBr}$	<b>119</b>
$\text{BeCr}_2\text{O}_7$	<b>224.9</b>
$\text{Co}(\text{ClO}_3)_2$	<b>225.8</b>
$\text{Cu}_2\text{C}_4\text{H}_4\text{O}_6$	<b>275.2</b>
$\text{CuSO}_4 \cdot 7 \text{H}_2\text{O}$	<b>285.7</b>

COMPOUND	FORMULA MASS
$\text{ZnCl}_2$	<b>136.3</b>
$\text{K}_3\text{PO}_4$	<b>212.3</b>
$\text{Al}_2(\text{SO}_4)_3$	<b>342.2</b>
$\text{MgCrO}_4$	<b>140.3</b>
$\text{CaC}_4\text{H}_4\text{O}_6$	<b>188.2</b>
$\text{NaCl}$	<b>58.4</b>
$\text{K}_2\text{Cr}_2\text{O}_7$	<b>294.2</b>
$\text{H}_2\text{SO}_4$	<b>98.1</b>
$\text{Cu}(\text{OH})_2$	<b>97.6</b>
$\text{MgSO}_4 \cdot 5 \text{H}_2\text{O}$	<b>210.4</b>