

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

Graded problems are worth 2 points each. This assignment is graded out of **18** points.

For each compound make two (2) drawings:

**Drawing a:** Draw the molecule formed using Lewis dot diagrams for each of the atoms. Circle the electron pairs that comprise covalent bonds.**Drawing b:** Draw the molecule again as a “ball and stick” diagram, replacing circled electron pairs with lines to indicate covalent bonds. You may rearrange the atoms to make the drawing neater.1) O<sub>2</sub>

a	b
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2) H<sub>2</sub>O

a	b
---	---

3) CH<sub>4</sub>

a	b
---	---

4) HCOOH

a	b
---	---

5) C<sub>2</sub>H<sub>5</sub>OH

a	b
---	---

6) **Graded.** NH<sub>3</sub>

a	b
---	---

7) **Graded.** SiO<sub>2</sub>

a	b
---	---

8) **Graded.**  $\text{SiCl}_4$ 

a	b
---	---

9) **Graded.**  $\text{HNO}$ 

a	b
---	---

10) **Graded.**  $\text{C}_2\text{F}_2$ 

a	b
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11) **Graded.**  $C_3H_7OH$ 

a	b
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12) **Graded.**  $C_3H_5OH$ 

a	b
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13) **Graded.**  $Si_2H_2O$ 

a	b
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14) **Graded.**  $H_3C_3OOH$ 

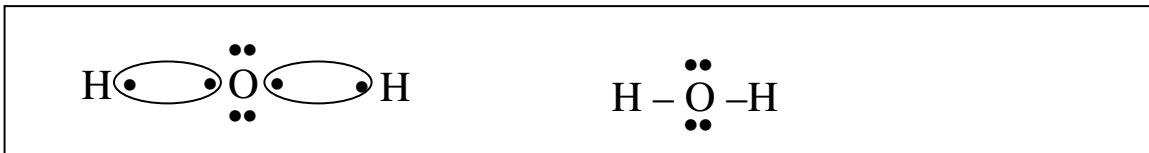
a	b
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## Selected Answers

1) O<sub>2</sub>



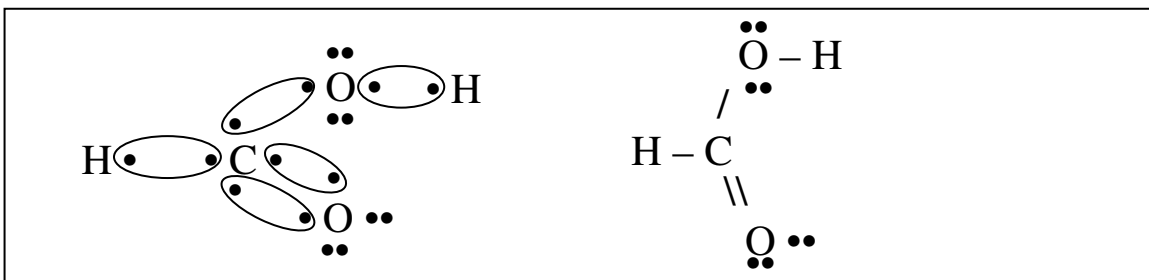
2) H<sub>2</sub>O



3) CH<sub>4</sub>



4) HCOOH



5) C<sub>2</sub>H<sub>5</sub>OH

